

More Mixed Ionic and Covalent Compound Naming

Name 'em:

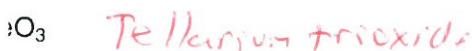
- 1) Br_2 Bromine
- 2) $(\text{NH}_4)_2\text{SO}_4$ Ammonium sulfate
- 3) HNO_2 Nitrous Acid
- 4) $\text{Pb}(\text{CO}_3)_2$ Pb(IV) Carbonate
- 5) N_2O dinitrogen Monoxide
- 6) ZnF_2 Zinc Fluoride
- 7) CCl_4 Carbon Tetrachloride
- 8) H_2S Hydro sulfide acid
- 9) $\text{Al}(\text{OH})_3$ Aluminum Hydroxide
- 10) K_2O Potassium Oxide

Formulas, please:

- 11) sulfur trioxide SO_3
- 12) tin(II) sulfide SnS
- 13) magnesium acetate $\text{Mg}(\text{C}_2\text{H}_3\text{O}_2)_2$
- 14) hydrophosphoric acid H_3P
- 15) cadmium chloride CdCl_2
- 16) iron(III) phosphate FePO_4
- 17) dinitrogen trioxide N_2O_3
- 18) fluorine F_2
- 19) vanadium(V) bicarbonate $\text{V}(\text{HCO}_3)_5$
- 20) ammonia NH_3

Mixed Ionic and Covalent Naming Practice

Given formulas provided, name each of the following compounds:



Name the following covalent compounds using the names shown:

